$$wt\% = \frac{m(std) \times M_W(\mathbf{3}) \times molar\ ratio \times P(std)}{m(\mathbf{3}) \times M_W(std)} = \mathbf{99}\%$$

m(**5**) = 19.3 mg; M_W(**5**) = 193.25 g/mol

m(std) = 16.6 mg; M_W(std) = 168.19 g/mol

P(std) = 98%; molar ratio = (3,08/3)/(3/3) = 1.03